

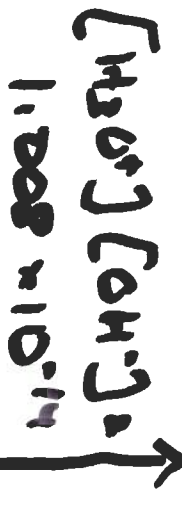
[Weak] Acid



$$pH = -\log [H_3O^+]$$

pH

Most common way to characterize acidity/basicity of aqueous solns
pH < 7 acidic
pH = 7 neutral
pH > 7 basic



$$pOH = -\log [OH^-]$$

$$pH + pOH = 13.9965$$

[strong] base



Conc. in Molarity

[weak] base

$$M = \frac{\text{mol}}{L} \rightarrow \text{grams}$$

Conc. values are regular numbers - normal rules for sig figs

pH & pOH are logarithms

sig figs = # places to right of decimal point